

AD-A100 779

THE ORBITAL-OVERLAP FACTOR IN ELECTRON TRANSFER
SENSITIVITY OF HOMOGENEOUS (U) PURDUE UNIV LAFAYETTE IN
DEPT OF CHEMISTRY R H NIELSON ET AL. 25 SEP 67 TR-69
N00014-86-K-0336

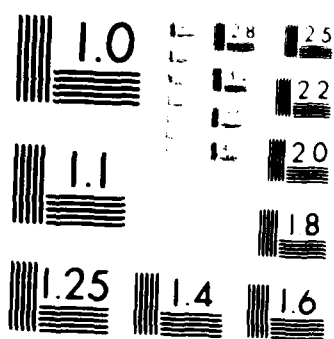
1/1

UNCLASSIFIED

F/G 7/4

NL





Resolution Test Chart
NBS 1010-A-10

4

DTIC FILE COPY

OFFICE OF NAVAL RESEARCH

Contract N00014-86-K-0556

Technical Report No. 69

The Orbital-Overlap Factor in Electron Transfer:
Sensitivity of Homogeneous Self-Exchange Kinetics for
Some Metallocenes to Electronic Structure

by

R. M. Nielson, M. N. Golovin, G. E. McManis, and M. J. Weaver

Prepared for Publication

in the

Journal of the American Chemical Society

Purdue University

Department of Chemistry

West Lafayette, Indiana 47907

DTIC
ELECTE
DEC 30 1987
S H D

September 25, 1987

Reproduction in whole, or in part, is permitted for any purpose of the United States Government.

* This document has been approved for public release and sale: its distribution is unlimited.

AD-A188 779

AD-A188 779

REPORT DOCUMENTATION PAGE

1a REPORT SECURITY CLASSIFICATION Unclassified		1b RESTRICTIVE MARKINGS	
2a SECURITY CLASSIFICATION AUTHORITY		3 DISTRIBUTION/AVAILABILITY OF REPORT Approved for public release and sale; its distribution is unlimited.	
2b DECLASSIFICATION/DOWNGRADING SCHEDULE		5. MONITORING ORGANIZATION REPORT NUMBER(S)	
4 PERFORMING ORGANIZATION REPORT NUMBER(S) Technical Report No. 69		7a NAME OF MONITORING ORGANIZATION Division of Sponsored Programs Purdue Research Foundation	
5a NAME OF PERFORMING ORGANIZATION Purdue University Department of Chemistry		7b. ADDRESS (City, State, and ZIP Code) Purdue University West Lafayette, Indiana 47907	
5b ADDRESS (City, State, and ZIP Code) Purdue University Department of Chemistry West Lafayette, Indiana 47907		9 PROCUREMENT INSTRUMENT IDENTIFICATION NUMBER Contract No. N00014-86-K-0556	
6a NAME OF FUNDING/SPONSORING ORGANIZATION Office of Naval Research		10 SOURCE OF FUNDING NUMBERS	
6b ADDRESS (City, State, and ZIP Code) 800 N. Quincy Street Arlington, VA 22217		PROGRAM ELEMENT NO	
		PROJECT NO	
		TASK NO	
		WORK UNIT ACCESSION NO	
7c TITLE (Include Security Classification) The Orbital-Overlap Factor in Electron Transfer: Sensitivity of Homogeneous Self-Exchange Kinetics for Some Metallocenes to Electronic Structure			
8 PERSONAL AUTHOR(S) R. M. Nielson, M. N. Golovin, G. E. McManis, and M. J. Weaver			
9a TYPE OF REPORT Technical		13b TIME COVERED FROM 10/1/86 TO 9/30/87	
		14 DATE OF REPORT (Year, Month, Day) September 25, 1987	
		15 PAGE COUNT	
16 SUPPLEMENTARY NOTATION			
COSATI CODES		18 SUBJECT TERMS (Continue on reverse if necessary and identify by block number)	
FIELD	GROUP	SUB-GROUP	
1			
2			
3			
4			
5			
6			
7			
8			
9			
10			
11			
12			
13			
14			
15			
16			
17			
18			
19			
20			
21			
22			
23			
24			
25			
26			
27			
28			
29			
30			
31			
32			
33			
34			
35			
36			
37			
38			
39			
40			
41			
42			
43			
44			
45			
46			
47			
48			
49			
50			
51			
52			
53			
54			
55			
56			
57			
58			
59			
60			
61			
62			
63			
64			
65			
66			
67			
68			
69			
70			
71			
72			
73			
74			
75			
76			
77			
78			
79			
80			
81			
82			
83			
84			
85			
86			
87			
88			
89			
90			
91			
92			
93			
94			
95			
96			
97			
98			
99			
100			

The Orbital-Overlap Factor in Electron Transfer:
Sensitivity of Homogeneous Self-Exchange Kinetics for Some Metallocenes
to Electronic Structure

Roger M. Nielson, M. Neal Golovin, George E. McManis, and Michael J. Weaver*

Department of Chemistry

Purdue University

West Lafayette, IN 47907



Accession For	
NTIS GRA&I	<input checked="checked" type="checkbox"/>
DTIC TAB	<input type="checkbox"/>
Unannounced	<input type="checkbox"/>
Justification	
By	
Distribution/	
Availability Codes	
Dist	Avail and/or Special
A-1	

Abstract

Rate constants, k_{ex}^h , for the electron self exchange of cobaltocenium-cobaltocene, $\text{Cp}_2\text{Co}^{+/0}$, and for the decamethyl derivative $(\text{Cp-Me}_5)_2\text{Co}^{+/0}$ obtained by using the NMR line broadening technique in acetonitrile and dimethylsulfoxide are compared with corresponding data for the ferrocene couples $\text{Cp}_2\text{Fe}^{+/0}$ and $(\text{Cp-Me}_5)_2\text{Fe}^{+/0}$ and bis(benzene)chromium (I)/(0). The rate constants in a given solvent display a marked sensitivity to the reactant structure, the k_{ex}^h values being about tenfold larger for $\text{Cp}_2\text{Co}^{+/0}$ relative to $\text{Cp}_2\text{Fe}^{+/0}$; decamethyl substitution yields tenfold increases in k_{ex}^h for both these couples. A relationship is established between these ca. 100 fold rate variations and the nature of the donor and acceptor orbitals. In particular, the markedly slower self-exchange kinetics observed for $\text{Cp}_2\text{Fe}^{+/0}$ relative to $\text{Cp}_2\text{Co}^{+/0}$ are consistent with the much greater ligand-delocalized character of the $4e_{1g}$ orbital involved in the latter electron transfer as compared with the $4e_2$ or $8a_{1g}$ orbital for the former reaction. The same argument is likely to account for the similar relative rate ratio for $(\text{Cp-Me}_5)_2\text{Fe}^{+/0}$ versus $(\text{Cp-Me}_5)_2\text{Co}^{+/0}$. These rate differences are very unlikely to be due to variations in nuclear reorganization factors since the molecular structure of these couples are virtually identical, and feature only small differences ($\leq 0.3 \text{ kcal mol}^{-1}$) in the inner-shell barriers. The results therefore provide clear evidence for the influence of donor-acceptor electronic coupling in outer-sphere redox reactivity.

We have recently been examining the electrochemical electron exchange kinetics of various metallocene redox couples as a function of the solvent in order to probe the possible role of solvent relaxation dynamics upon the barrier-crossing frequencies.¹ The couples selected include cobaltocenium-cobaltocene ($\text{Cp}_2\text{Co}^{+/0}$ where Cp = cyclopentadiene), and the decamethyl derivatives ($\text{Cp-Me}_5)_2\text{Co}^{+/0}$ (where Cp-Me₅ = pentamethylcyclopentadiene) and ($\text{Cp-Me}_5)_2\text{Fe}^{+/0}$.^{1b} These as well as other structurally related metallocene and metal arene couples yield similar electrochemical reactivities in a given solvent, the small (ca. twofold or less) rate variations being consistent with the minor differences in the inner-shell (i.e. reactant bond distortional) barrier, ΔG_{is}^* , anticipated on the basis of structural data.^{1b} A major virtue of these systems as model reactants for solvent dynamical studies is that the electron-transfer barrier is dominated by the reorganization of the surrounding solvent.^{1b} In addition, the metallocenes approach the apparent spherical geometry desirable for theoretical comparisons.

Most recently, we have been expanding these studies to include measurements of the corresponding self-exchange kinetics in homogeneous solution by utilizing the NMR line-broadening technique. A primary objective is to compare the form of the solvent-dependent kinetics at electrochemical interfaces and in homogeneous solution in order to ascertain if the nature and extent of the solvent dynamical effects is different in these two redox environments. During these studies it became clear that, in contrast to the electrochemical reactions, the homogeneous-phase systems display a notable sensitivity to the reactant electronic structure.

In the present communication we report some pertinent results, for $\text{Cp}_2\text{Co}^{+/0}$ and $(\text{Cp-Me}_5)_2\text{Co}^{+/0}$ self-exchange in acetonitrile and dimethylsulfoxide, which along with corresponding published data for $\text{Cp}_2\text{Fe}^{+/0}$, $(\text{Cp-Me}_5)_2\text{Fe}^{+/0}$,³ and for bis(benzene)Cr(I)/(0) $[(\text{C}_6\text{H}_6)_2\text{Cr}^{+/0}]$,⁴ illustrate the dependence of the rate constants in a given solvent upon the nature of the metal as well as the ligand structure. A correlation is established here between the ca. 100 fold rate variations for these metallocene couples in a given solvent and the character of the orbitals involved in the electron transfer. The results implicate the importance of donor-acceptor orbital overlap to electron-transfer reactivity for such simple homogeneous-phase outer-sphere processes.

Experimental

Cobaltocene was obtained from Strem Chemicals. Cobaltocenium tetrafluoroborate was prepared by oxidation of Cp_2Co with tetrafluoroboric acid (Alfa); the corresponding decamethyl derivatives were synthesized as described in ref. 5. Acetonitrile and dimethylsulfoxide were "high purity" grade from Burdick and Jackson; the former was purified further by distillation over phosphorus pentoxide. Deuterated solvents were from Aldrich. All solutions for NMR measurements were prepared in 5mm tubes in a nitrogen-filled glove box. The $\text{Cp}_2\text{Co}^{+/0}$ system employed oxidized/reduced form concentrations of 0.02-0.12M / 0.7-20mM; for the $(\text{Cp-Me}_5)_2\text{Co}^{+/0}$ system these were 0.01-0.025M / 0.8-8mM. The solutions also contained 1% tetramethylsilane as a chemical shift reference.

Proton NMR spectra were collected using Nicolet NT 200 and NT 470 instruments. For data collection in protiated solvents, the interfering solvent proton peaks were diminished by homonuclear irradiation of

undesirable resonances. The rate constants obtained in the protiated and deuterated solvents generally agreed to within 10%. The NMR chemical shifts for the oxidized/reduced form mixtures equalled the weighted averages of those for the individual redox states measured separately. This indicates that the electron-transfer process conforms to the "fast exchange" region; the procedure of Wahl et al.⁶ was used to extract the desired rate constants for homogeneous self exchange, k_{ex}^{h} , by employing the measured chemical shifts and line widths of solutions of the oxidized/reduced mixture and of the separate redox states.⁷

Results and Discussion

Table I summarizes rate constants for homogeneous self exchange, k_{ex}^{h} , as determined here for the cobaltocene redox couples $\text{Cp}_2\text{Co}^{+/\circ}$ and $(\text{Cp-Me}_5)_2\text{Co}^{+/\circ}$ in acetonitrile and dimethylsulfoxide at 25 (± 0.5)°C. These values were obtained for solutions containing 0.1M tetraethylammonium tetrafluoroborate; k_{ex}^{h} values only about 10% larger were obtained in the absence of this added electrolyte. (Pertinent information regarding the nmr data are summarized in footnote 7.) Table I also includes corresponding k_{ex}^{h} values for $\text{Cp}_2\text{Fe}^{+/\circ}$, $(\text{Cp-Me}_5)_2\text{Fe}^{+/\circ}$, and $(\text{C}_6\text{H}_6)_2\text{Cr}^{+/\circ}$, taken from refs. 3 and 4, which were obtained using proton NMR and ESR line broadening techniques, respectively. The former values were also replicated in our laboratory.²

Comparison of these rate constants in a given solvent reveals that the cobaltocenium-cobaltocene couple exhibits substantially (ca. tenfold) higher k_{ex}^{h} values than the ferrocenium-ferrocene couple. Comparable (ca. tenfold) rate increases are also seen upon methylating the cyclopentadienyl rings for both the cobalt and iron systems. The $(\text{Cp-Me}_5)_2\text{Co}^{+/\circ}$ couple

exhibits k_{ex}^h values that are about 100-fold larger than for $\text{Cp}_2\text{Fe}^{+/0}$, with both $\text{Cp}_2\text{Co}^{+/0}$ and $(\text{Cp-Me}_5)_2\text{Fe}^{+/0}$ yielding similarly intermediate k_{ex}^h values in a given solvent (Table I). Roughly comparable rate ratios for these four reactions are also observed in a number of other solvents;² acetonitrile and dimethylsulfoxide are chosen here chiefly for illustrative purposes. The value of k_{ex}^h for $(\text{C}_6\text{H}_5)_2\text{Cr}^{+/0}$ in dimethylsulfoxide, $6 \times 10^7 \text{ M}^{-1} \text{ s}^{-1}$, is relatively high, being intermediate between that for $\text{Cp}_2\text{Co}^{+/0}$ and $(\text{Cp-Me}_5)_2\text{Co}^{+/0}$.

In order to elucidate the origin of these surprising rate differences, it is useful to express the rate constant as^{1,8}

$$k_{\text{ex}}^h = K_p \kappa_{\text{el}} \nu_n \exp [-(\Delta G_{\text{os}}^* + \Delta G_{\text{is}}^*)/RT] \quad (1)$$

where K_p is an equilibrium constant for forming the precursor state from the separated reactants, ν_n is the nuclear frequency factor, and κ_{el} is the electronic transmission coefficient (i.e. the probability that electron tunneling will occur upon formation of the nuclear transition state). The form of Eq. (1) applies to both homogeneous-phase and electrochemical outer-sphere processes.⁹

As stated above, the inner-shell barriers, ΔG_{is}^* , for the present reactants are small. These values can readily be estimated from bond distance and vibrational spectroscopic data, using⁸

$$\Delta G_{\text{is}}^* = 0.5 n f_{\text{is}} (\Delta a/2)^2 \quad (2)$$

where n is the number of bonds undergoing distortion, Δa is the change in

bond distance between the oxidized and reduced forms, and f_{is}^r is the "reduced" force constant for each bond obtained from the individual force constants for the oxidized and reduced species using $f_{is}^r = 2 f_{is}^{ox} f_{is}^{red} / (f_{is}^{ox} + f_{is}^{red})$. The last can be obtained from

$$f_{is} = 4\pi^2 \nu_{is}^2 \mu \quad (2a)$$

where ν_{is} is the observed vibrational frequency (s^{-1}) and μ is the reduced mass.

Fortunately, accurate bond-distance data are available for the present redox couples which indicate that the only significant structural difference between the oxidized and reduced forms is a small increase (or decrease) of the metal-ring distance. These Δa values are⁹: $(Cp-Me_5)_2Co^{+/0}$, 0.045Å; $Cp_2Co^{+/0}$, \approx 0.05Å; $(Cp-Me_5)_2Fe^{+/0}$, -0.035Å; $Cp_2Fe^{+/0}$, ca. -0.035Å;¹⁰ $(C_6H_6)_2Cr^{+/0}$, \approx 0.045Å. Since the inner-shell distortion for these couples involves a symmetric metal-ring vibration, the appropriate ν_{is} values required are those for the Raman-active A_{1g} stretching mode. Such data are available for $Cp_2Co^{+/0}$,^{12,13} $Cp_2Fe^{+/0}$,¹³ and $(C_6H_6)_2Cr^{+/0}$.¹⁴ The resulting force constants¹⁵ together with the above Δa values inserted into Eq. (2), with $n = 4$ (i.e. two vibrating bonds for each reactant), yield ΔG_{is}^* values for these self-exchange reactions of ca. 0.7, 0.35, and 0.5 kcal mol⁻¹, respectively. While these estimates of ΔG_{is}^* cannot be regarded as precise, they are likely correct to about 20-30%, i.e. within ca. 0.2 kcal mol⁻¹. Although the force constants for the $(Cp-Me_5)_2Co^{+/0}$ and $(Cp-Me_5)_2Fe^{+/0}$ couples are known with less certainty,^{17,18} it is likely that similar or slightly smaller ΔG_{is}^* values

apply to these systems.

These small variations in ΔG_{is}^* between the self-exchange process considered here eliminate this factor as a plausible explanation for the observed rate differences. In particular, the tenfold greater rate constants for $Cp_2Co^{+/0}$ with respect to $Cp_2Fe^{+/0}$ cannot be accounted for by differences in ΔG_{is}^* since this term is slightly ($0.2-0.3 \text{ kcal mol}^{-1}$) larger for the former couple. (Note that to account for the observed tenfold difference in k_{ex}^h on this basis would require that ΔG_{is}^* is about $1.4 \text{ kcal mol}^{-1}$ smaller for $Cp_2Co^{+/0}$.)

In view of their almost identical size and structure, the outer-shell reorganization energy, ΔG_{os}^* , for $Cp_2Co^{+/0}$ and $Cp_2Fe^{+/0}$ should be virtually identical, although ΔG_{os}^* may be slightly smaller for the decamethyl derivatives and $(C_6H_6)_2Cr^{+/0}$ due to their larger effective radius. It therefore might be argued that the rate increases seen upon methylation are due at least in part to a decrease in the outer-shell barrier, as well as to slightly smaller ΔG_{is}^* values. This possibility, however, appears unlikely since the rate constants for $Cp_2Co^{+/0}$, $(Cp-Me_5)_2Co^{+/0}$, and $(C_6H_6)_2Cr^{+/0}$ electrochemical exchange in a given solvent are almost identical, even though the ΔG_{os}^* values should be comparable to those for the corresponding homogeneous reactions.^{1b}

Consequently, the primary origin of the sizable differences in k_{ex}^h for these five redox couples lies almost undoubtedly in corresponding variations in the preexponential factor $K_p \kappa_{el} \nu_n$. The nuclear frequency factor ν_n is expected to be dominated largely by solvent reorganization dynamics rather than by inner-shell vibrations in view of the predominant contribution of ΔG_{os}^* to the electron-transfer barrier.^{1b,c} We have

recently been utilizing such metallocene systems to test experimentally the theoretical prediction¹² that ν_n is determined by "overlapped" solvent motion in most solvents. Although interesting deviations from this prediction have come to light,^{10,2} it suffices to note here that the solvent dependence of the rate constants for homogeneous self-exchange as well as electrochemical exchange for the present metallocene couples are roughly consistent with such theoretical expectations.^{1,2} In a given solvent, therefore, ν_n should be very similar for all the present reactions.

This leaves variations in K_p and/or κ_{e1} as constituting the only tenable explanation for the observed variations in k_{ex}^h with reactant structure. For such homogeneous-phase outer-sphere processes, K_p will be determined by electrostatic interactions between the reactant pair ("work terms") together with the statistical probability of forming such a pair from the (presumed spherical) reactants. The former component should be absent here since one reactant is uncharged. In this case a suitable expression for K_p is^{8,22}

$$K_p = 4\pi N r^2 \delta r \quad (2)$$

where N is Avogadro's number, r is the average separation of the reacting centers in the transition state, and δr is the effective "reaction zone thickness", i.e. the range of internuclear separations larger than that corresponding to molecular "contact" which contribute substantially to the overall reaction rate. The last term should be sensitive to the effective distance over which efficient electron tunneling between the donor and acceptor orbitals can occur, and therefore is closely linked to the corresponding value of κ_{e1} in Eq. (1).

It is useful here to identify two distinct limiting cases. Firstly case 1: the electronic coupling between the reaction centers may be sufficient so that efficient electron tunneling occurs (i.e. reaction adiabaticity is achieved, not only upon reactant contact, but also for significantly larger internuclear separations. In this case $\kappa_{e1} \sim 1$ and the magnitude of δr reflects the range of larger internuclear separations over which reaction adiabaticity is maintained. Alternatively (case 2), the reaction may be substantially nonadiabatic even at the reactant contact distance where effectively $\kappa_{e1} \ll 1$ and δr will be small, ca. 0.5 \AA .^{22,23} Additionally, κ_{e1} may be sensitive to the reactant pair geometry, so that only certain specific internuclear configurations will correspond to sufficient electronic coupling so to yield viable reaction channels (vide infra). In this case, therefore, the form of the K_p expression, and hence the magnitude of K_p as well as κ_{e1} can be very sensitive to the donor-acceptor electronic coupling.

There is ample evidence from ab initio electronic structural and other calculations that outer-sphere electron transfer even between small molecules tends to border on case 2 rather than case 1 behavior.²⁴ Some indirect experimental evidence based on absolute theory-experiment comparisons of rate parameters^{8,25} and from analyses of ligand effects upon the rates of some related reactions²⁶ are also in harmony with this assertion. It is therefore of interest to compare the electronic structures of the metallocenes considered here to ascertain if the observed rate differences can be accounted for on this basis.

The primary factor influencing the degree of electronic coupling and hence $K_p \kappa_{e1}$ is the extent of spatial overlap between the acceptor and donor

orbitals, as this determines the magnitude of the electronic matrix coupling element, H_{12} .²⁴ This overlap for the present systems will tend to be greater for orbitals having a greater delocalization over the Cp rings.

Information on this question can be extracted from a number of ab initio, $X\alpha$ and semiempirical theoretical studies of ferrocene²⁷ and cobaltocene.²⁸ The effective donor orbital will be the HOMO for the reduced form of the couple. This orbital for ferrocene appears to be extremely metal centered. Thus INDO-SCF calculations identify it as having $4e_2$ symmetry with 86% metal character, with a $8a_{1g}$ orbital having 90% metal character lying immediately below it in energy;^{27a} $X\alpha$ calculations identify the HOMO as the latter orbital.^{27c} This predicted metal-centered electron distribution is consistent with ESR estimates for Cp_2Fe^+ .²⁹ The highly metal-localized nature of the HOMO for ferrocene is strikingly apparent when viewed in pictorial form.³⁰ For cobaltocene, by comparison, the HOMO is distinctly more ligand centered and therefore delocalized; $X\alpha$ ^{28a} and LCAO-HFS^{28b} calculations identify this as the singly-occupied $4e_{1g}$ orbital with only around 50-55% metal character, in harmony with ESR observations for cobaltocene.³¹

Since the lowest unfilled molecular orbital (LUMO) on the oxidized reactants, ferrocenium and cobaltocenium, correspond to the same orbital assignment as on the corresponding reduced species,^{27a,b,28} we deduce that the extent of donor-acceptor orbital overlap for the $Cp_2Co^+-Cp_2Co$ reaction will be greater than for the $Cp_2Fe^+-Cp_2Fe$ exchange within a given internuclear geometry. A complication, however, arises from the relaxation in electron distribution which occurs for these systems upon removal of an electron from the HOMO; this reflects the multielectron screening of the

formal metal charge and leads to relatively small and similar increases in the net metal charge upon the oxidation of both Cp_2Fe and Cp_2Co .^{27a,28a} The implications of this "charge relaxation effect" upon the degree of electronic coupling have recently been addressed.^{24a} However, this effect is expected only to modify rather than to eliminate the enhancement of H_{12} and hence $K_{p\kappa_{el}}$ that is predicted for $\text{Cp}_2\text{Co}^+-\text{Cp}_2\text{Co}$ relative to $\text{Cp}_2\text{Fe}^+-\text{Cp}_2\text{Fe}$ self exchange. The substantially larger k_{ex}^h values for the former reaction can therefore be understood at least qualitatively on this basis.

The effect of methylation of the Cp rings upon k_{ex}^h for these reactions is also consistent with such electronic considerations. Although detailed theoretical presentations involving individual orbitals are unavailable, LCAO-SCF calculations predict that the net electron density on the Cp ring is increased markedly by methylation.³² A similar deduction has been made on the basis of NMR data for cobaltocenium^{33a} and ferrocene,^{33b} and also from photoelectron spectra.³⁴ Although it is difficult to infer from such information the desired effect upon the orbital involved in electron transfer, it nonetheless seems reasonable to anticipate a larger donor-acceptor orbital overlap resulting from ring methylation.

The relative k_{ex}^h values for the four metallocene self-exchange reactions considered here are therefore nicely consistent with electronic overlap considerations, especially for $\text{Cp}_2\text{Co}^{+/0}$ versus $\text{Cp}_2\text{Fe}^{+/0}$ for which detailed electronic structures are available. Given the metal-centered nature of the redox orbitals for $\text{Cp}_2\text{Fe}^{+/0}$, one can envisage that the internuclear geometries leading to reaction would be restricted to those with the reactants approaching along a common five-fold axis, where some direct metal-metal orbital overlap is most likely to be achieved. As noted

above, such a "geometric" limitation on the electron-transfer rate via the extent of orbital overlap and hence upon κ_{el} can be viewed as arising from correspondingly smaller K_p values than expected from the spherical reactant model [i.e. from Eq. (2)]. Compared to $\text{Cp}_2\text{Co}^{+/0}$, therefore, the tenfold smaller k_{ex}^h value for $\text{Cp}_2\text{Fe}^{+/0}$ can be viewed as arising from the smaller probability of forming internuclear reactant configurations with sufficient donor-acceptor orbital overlap to contribute importantly to electron transfer.

The comparison of k_{ex}^h for these two reactions with that for $(\text{C}_6\text{H}_6)_2\text{Cr}^{+/0}$ is also of interest. Semiempirical and $X\alpha$ calculations²⁴ indicate that the HOMO of $(\text{C}_6\text{H}_6)_2\text{Cr}$ is a metal-localized $8a_{1g}$ orbital, although a substantially delocalized $4e_{2g}$ orbital lies close in energy and has been assigned as the HOMO on the basis of CNDO and SCF calculations.³⁵ However, given the relatively rapid self exchange of $(\text{C}_6\text{H}_6)_2\text{Cr}^{+/0}$, the k_{ex}^h being intermediate between $\text{Cp}_2\text{Co}^{+/0}$ and $(\text{Cp-Me}_5)_2\text{Co}^{+/0}$ (Table I), one would expect that substantial donor-acceptor orbital overlap is nevertheless achieved. Indeed, on the basis of INDO-SCF calculations, $4e_{2g}$ orbitals is predicted to be only slightly lower in energy than the ia_{1g} for $(\text{C}_6\text{H}_6)_2\text{Cr}^+$; moreover, these orbitals are increasingly degenerate for the larger Cr-ring bond distances³⁷ that would be appropriate for the Cr(I)/(0) transition state. The likely involvement of the ligand-delocalized $4e_{2g}$ orbital in the electron transfer may therefore account for the facile kinetics observed for this reaction.

In addition to examining the variations in k_{ex}^h with metallocene structure in comparison with the theoretical expectations it is also informative to explore experiment-theory comparisons for the individual

rate constants themselves.^{8,25} Table I contains such theoretical estimates of k_{ex}^{h} , labelled $k_{\text{ex}}^{\text{th}}$, for each of the metallocene couples and $(\text{C}_6\text{H}_6)_2\text{Cr}^{+/0}$ in acetonitrile and dimethylsulfoxide. These estimates were calculated essentially as described in ref. 1b, using Eq. (1). (See the Table footnotes and Table V of ref. 1b for details.) Thus the ν_{n} values were obtained from the "overdamped" solvent relaxation model using Eq. (9) of ref. 1b. (Note that the numerical value of ν_{n} should be relatively insensitive to the extent of donor-acceptor electronic coupling.³⁸) The reactant radius is taken as 3.8 Å in each case, and K_{p} is calculated from Eq. (2) with $r = 7.6$ Å (i.e. twice the reactant radius) and $\delta r = 0.6$ Å, yielding $K_{\text{p}} = 0.26 \text{ M}^{-1}$. The small difference between the $k_{\text{ex}}^{\text{th}}$ values in Table I for the different systems in a given solvent arise from the slightly different estimates of ΔG_{is}^* (vide supra).

Comparison between the corresponding values of k_{ex}^{h} and $k_{\text{ex}}^{\text{th}}$ in Table I shows that the experimental values are significantly smaller than the theoretical estimates only for $\text{Cp}_2\text{Fe}^{+/0}$. This finding that $k_{\text{ex}}^{\text{h}} < k_{\text{ex}}^{\text{th}}$ is suggestive of nonadiabatic (i.e. case 2) behavior for this system since the latter values are derived by presuming that $\kappa_{\text{el}} = 1$, (i.e. that adiabaticity is achieved, albeit with a small "reaction zone thickness", $\delta r = 0.6$ Å). The comparable or larger values of k_{ex}^{h} relative to $k_{\text{ex}}^{\text{th}}$ observed for the other metallocene reactions (Table I) might be inferred as signaling the onset of reaction adiabaticity for these systems. If this is so, this would not be consistent with the above assertion that the observed rate variations are due primarily to electronic coupling factors. However, the observation that $k_{\text{ex}}^{\text{h}} > k_{\text{ex}}^{\text{th}}$ for $(\text{Cp-Me}_5)_2\text{Co}^{+/0}$ is more likely to be due to systematic uncertainties in the theoretical estimates of ν_{n} , ΔG_{os}^* and/or

K_p . (This issue will be addressed in some detail elsewhere.²⁾

As noted above, in contrast to these homogeneous self-exchange processes the rate constants for the corresponding electrochemical exchange reactions at mercury electrodes display virtually no sensitivity to the metallocene structure.^{1b,39} The simplest interpretation of this disparate behavior is that the electrochemical reactions involve sufficiently strong electronic coupling so that adiabaticity is achieved in each case, i.e. "case 1" behavior is obtained. There are some indications from other measurements that electrochemical reactions can be markedly nonadiabatic, at least for surface - reaction site distances greater than ca. 6-8Å.^{23,40} On the other hand, there is evidence that simple outer-sphere electrochemical processes involving small metal complexes in aqueous media are more adiabatic than for the corresponding homogeneous-phase reactions.²⁵ Although the precursor-state geometries for such outer-sphere electrochemical processes are not known precisely, it is not unreasonable to envisage the reactant to lie suitably close to the surface so that adiabaticity is achieved. This would especially be the case if the reactant is able to at least partly penetrate the "inner layer" of solvent molecules adjacent to the metal surface.

Despite such uncertainties, the present results provide unusually direct evidence for the importance of orbital-overlap factors in outer-sphere redox reactivity. The differences in k_{ex}^h between the $Cp_2Fe^{+/0}$ and $Cp_2Co^{+/0}$, and $(Cp-Me_5)_2Fe^{+/0}$ and $(Cp-Me_5)_2Co^{+/0}$ self-exchange reactions are of particular significance since these pairs of redox couples have structural properties that are otherwise virtually identical. Although such effects may well be prevalent in many other systems, they usually

would remain masked by the presence of other obfuscating factors, such as large unknown variations in inner-shell barriers, work terms, and so on, when the kinetics of related reactions are compared. It would be worthwhile to evaluate k_{ex}^h for other metallocene or arene couples featuring substituents that exert large electronic perturbations on the aromatic rings. The quantitative calculation of electronic matrix coupling elements for these systems using ab initio methods²⁴ would also be of considerable interest.

Acknowledgments

Some Raman spectroscopic data were obtained by Lam Leung and Dave Gosztola. This work is supported by the Office of Naval Research. The NMR instruments used in this work are supported by National Institute of Health Grant RR 01077 at Purdue. G.E.M. gratefully acknowledges financial support from a Department of Navy Graduate Research Fellowship.

References and Notes

1. (a) Weaver, M. J.; and Gennett, T., Chem. Phys. Lett., 1985, 113, 213; (b) Gennett, T.; Milner, D. F.; and Weaver, M. J., J. Phys. Chem., 1985, 89, 2787; (c) McManis, G. E.; Golovin, M. N.; and Weaver, M. J., J. Phys. Chem., 1986, 90, 6563.
2. Nielson, R. M.; McManis, G. E.; Golovin, M. N.; and Weaver, M. J., in preparation.
3. Yang, S. E.; Chan, M-S.; and Wahl, A. C., J. Phys. Chem., 1980, 84, 3094.
4. Li, T. T-T.; Weaver, M. J.; and Brubaker, C. H., Jr., J. Am. Chem. Soc., 1982, 104, 2381; Li, T. T-T.; and Brubaker, C. H., Jr., J. Organomet. Chem., 1981, 216, 223.
5. Kölle, U.; and Khouzami, F., Chem. Ber., 1981, 114, 2929.
6. Chan, M-S.; and Wahl, A. C., J. Phys. Chem., 1978, 24, 2542.
7. Typical experimental parameters obtained from nmr spectra used to

obtain rate constants for $\text{Cp}_2\text{Co}^{+/0}$ and $\text{Cp}_2'\text{Co}^{+/0}$ self exchange in Table I using the analysis procedure in ref. 6 are as follows (using nomenclature of Wahl et al.,⁶ as noted below). $\text{Cp}_2\text{Co}^{+/0}$ in acetonitrile: w_{DP} , 193 Hz; w_{P} , 119 Hz; w_{D} , 1 Hz; $\partial\nu$, -11228 Hz; c , 41.0 mM; f_{D} , 0.785. $\text{Cp}_2\text{Co}^{+/0}$ in DMSO: w_{DP} , 240 Hz; w_{P} , 152 Hz, w_{D} , 1 Hz; $\partial\nu$, 19981 Hz; c , 14.3 mM; f_{D} , 0.884. Parameters defined as: w_{DP} , w_{P} , w_{D} are peak widths of diamagnetic-paramagnetic mixture, and pure paramagnetic, diamagnetic species, respectively; $\partial\nu$ is "contact shift", c is total concentration (diamagnetic + paramagnetic species), f_{D} is mole fraction of diamagnetic species. The measured parameters for $\text{Cp}_2\text{Co}^{+/0}$ and $(\text{Cp-Me}_5)_2\text{Co}^{+/0}$ refer to fields of 200.07 and 469.58 MHz, respectively.

8. For a recent review, see Sutin, N., Prog. Inorg. Chem., 1983, 30, 441.
9. See Table I of ref. 1b for detailed bond distances and literature sources.
10. Although the difference in metal-ring bond lengths, Δa , between Cp_2Fe and Cp_2Fe^+ has not been evaluated since the latter crystal structure is apparently unavailable, the Δa value for $\text{Cp}_2\text{Fe}^{+/0}$ is almost certainly similar to that for $(\text{Cp-Me}_5)_2\text{Fe}^{+/0}$ in view of the identical metal-ring distance for Cp_2Fe and $(\text{Cp-Me}_5)_2\text{Fe}^{+/0}$.¹¹
11. Haaland, A., Acc. Chem. Res., 1979, 12, 415.
12. Hartley, D.; and Ware, M. J., J. Chem. Soc. A, 1969, 138.
13. Gachter, B. F.; Koningstein, J. A.; and Aleksanjan, V. T., J. Chem. Phys., 1975, 62, 4628.
14. Fritz, H. P.; Luttkie, W.; Stammreich, H.; and Forneris, R., Spectrochim. Acta., 1961, 17, 1068.
15. The $\nu_{\text{is}}^{\text{ox}}$ and $\nu_{\text{is}}^{\text{red}}$ values for these redox couples are taken as: $\text{Cp}_2\text{Co}^{+/0}$, 325,¹² 320 cm^{-1} ;¹³ $\text{Cp}_2\text{Fe}^{+/0}$, 310, 315 cm^{-1} ;¹³ $(\text{C}_6\text{H}_5)_2\text{Cr}^{+/0}$, 280, 270 cm^{-1} .¹⁴ Since the central metal atom remains stationary during the A_{1g} vibration, the effective force constant for both metal-ring vibrations can be obtained by inserting these values into Eq. (2a) with μ taken as the mass of a single ring; this yields f_{is}^r values of 4.0×10^5 , 3.75×10^5 , and 3.5×10^5 dyne. cm^{-1} , respectively.

16. These ΔG_{is}^* values for the homogeneous self-exchange reactions obtained here are roughly twice those calculated for the corresponding electrochemical exchange processes examined in ref. 1b. Besides the use of slightly different force constants here, this arises since a pair of reactants are involved in the former rather than the single reactant species as in the latter process.
17. For the $(Cp-Me_5)_2Fe^{+/0}$ couple, ν_{is}^{ox} and ν_{is}^{red} apparently occur at about 170 cm^{-1} .¹⁸ Assuming, as for the $Cp_2Fe^{+/0}$ couple that the ring can be treated as a single vibrating unit leads to a force constant $f_r \approx 2.3 \times 10^5\text{ dyne cm}^{-1}$. The somewhat smaller force constants derived for the decamethyl derivative are somewhat surprising, however, since almost identical values are anticipated from some theoretical considerations.¹⁹ Attempts to evaluate ν_{is}^{ox} and ν_{is}^{red} for $(Cp-Me_5)_2Co^+$ and $(Cp-Me_5)_2Co$ were thwarted by fluorescence and compound air instability, respectively.²⁰
18. Duggan, D. M.; and Hendrickson, D. N., *Inorg. Chem.*, 1975, 14, 955.
19. Luttile, H. P.; Ammeter, J. H.; Almlöf, J.; and Faegri, K., Jr., *J. Chem. Phys.*, 1982, 77, 2002.
20. Leung, L.-W. H.; and Gosztola, D., unpublished results.
21. For example, Calef, D. F.; and Wolynes, P. G., *J. Phys. Chem.*, 1983, 87, 3387.
22. Hupp, J. T.; and Weaver, M. J., *J. Electroanal. Chem.*, 1983, 152, 1.
23. Hupp, J. T.; and Weaver, M. J., *J. Phys. Chem.*, 1984, 88, 1463.
24. (a) Newton, M. D., *J. Phys. Chem.*, 1986, 90, 3734; (b) Newton, M. D., *ACS Symp. Ser.*, 1982, 198, 255.
25. Hupp, J. T.; and Weaver, M. J., *J. Phys. Chem.*, 1985, 89, 2795.
26. Endicott, J. F.; and Ramasami, T., *J. Phys. Chem.*, 1986, 90, 3740.
27. (a) Zerner, M. C.; Loew, G. H.; Kirchner, R. F.; and Mueller-Westerhoff, U. T., *J. Am. Chem. Soc.*, 1980, 102, 589; (b) Bagus, P. S.; Wahlgren, U. I.; and Almlöf, J., *J. Chem. Phys.*, 1976, 64, 2324; (c) Rosch, N.; and Johnson, K. H., *Chem. Phys. Lett.*, 1974, 24, 179; (d) Coutière, M.-M.; Demuynck, J.; and Veillard, A., *Theor. Chim. Acta*, 1972, 27, 281.

28. (a) Weber, J.; Goursot, A.; Penigault, E.; Ammeter, J. H.; and Bachmann, J., *J. Am. Chem. Soc.*, 1982, 104, 1491; (b) Famiglietti, C.; and Baerends, E. J., *Chem. Phys.*, 1981, 62, 407.
29. Prins, R., *Mol. Phys.*, 1970, 19, 603.
30. Hout, R. F., Jr.; Pietro, W. J.; Hehre, W. J., "A Pictorial Approach to Molecular Structure and Reactivity", Wiley, New York, 1984, Chapter 4.
31. Ammeter, J. H.; and Swalen, J. D., *J. Chem. Phys.*, 1972, 57, 678.
32. Luthi, H. P.; Ammeter, J. H.; Almlöf, J.; and Faegri, K., Jr., *J. Chem. Phys.*, 1982, 77, 2002.
33. (a) Materikova, R. B.; Babin, V. N.; Lyatifov, I. R.; Salimov, R. M.; Fedin, E. I.; and Petrovskii, P. V., *J. Organomet. Chem.*, 1981, 219, 259; (b) Materikova, R. B.; Babin, V. N.; Lyatifov, I. R.; Kurbanov, T. Kh.; Fedin, E. I.; Petrovskii, P. V.; and Lutsenko, A. I., *J. Organomet. Chem.*, 1977, 142, 81.
34. (a) Gassman, P. G.; Macomber, D. W.; and Hershberger, J. W., *Organometallics*, 1983, 2, 1470; (b) Cauletti, C.; Green, J. C.; Kelly, M. R.; Powell, P.; Van Tilborg, J.; Robbins, J.; and Smart, J., *J. Electron. Spect. Rel. Phen.*, 1980, 19, 327.
35. (a) Weber, J.; Geoffray, M.; Goursot, A.; and Penigault, E., *J. Am. Chem. Soc.*, 1978, 100, 3995; (b) Böhm, M. C., *Z. Naturforsch.*, 1982, 37a, 1193.
36. (a) Fitzpatrick, N. J.; Savariault, J-M.; Labarre, J-F. R., *J. Organomet. Chem.*, 1977, 127, 325; (b) Guest, M. F.; Hillier, I. H.; Higginson, B. R.; and Lloyd, D. R., *Mol. Phys.*, 1975, 29, 113.
37. Clack, D. W., and Smith, W., *Inorg. Chim. Acta*, 1976, 20, 93.
38. McManis, G. E.; Mishra, A. K.; and Weaver, M. J., *J. Chem. Phys.*, 1987, 86, 5550.
39. Admittedly, the electrochemical kinetics of $\text{Cp}_2\text{Fe}^{+/0}$ exchange at mercury can only be examined in a few solvents, such as acetonitrile, since the formal potential for this couple is close to, or positive of that for, mercury dissolution in most media. However, measurement of the exchange rate constant for this couple in acetonitrile under conditions as described in ref. 1c yielded a value, ca. 5 cm s^{-1} , similar to that obtained for $\text{Cp}_2\text{Co}^{+/0}$.^{1c}
40. (a) Li, T. T-T.; and Weaver, M. J., *J. Am. Chem. Soc.*, 1984, 106, 6106; (b) Weaver, M. J.; and Li, T. T-T., *J. Phys. Chem.*, 1986, 90, 3823.

TABLE I. Rate Constants for Homogeneous Self Exchange and Electrochemical Exchange of Various Metallocenes and Arene Redox Couples in Acetonitrile and Dimethylsulfoxide at 25°C.

Redox Couple ^a	k_{ex}^{h} ^b	k_{ex}^{h} ^b	$k_{\text{ex}}^{\text{thg}}$	$k_{\text{ex}}^{\text{thg}}$
	$\text{M}^{-1}\text{s}^{-1}$ ACN	$\text{M}^{-1}\text{s}^{-1}$ DMSO	$\text{M}^{-1}\text{s}^{-1}$ ACN	$\text{M}^{-1}\text{s}^{-1}$ DMSO
(Cp-Me ₅) ₂ Co ^{+/0}	4.3 (\pm 0.1) $\times 10^8$ ^c	1.8 (\pm 0.3) $\times 10^8$ ^c	3.3 $\times 10^7$	1.3 $\times 10^7$
Cp ₂ Co ^{+/0}	3.8 (\pm 0.2) $\times 10^7$ ^c	2.4 (\pm 0.3) $\times 10^7$ ^c	2.4 $\times 10^7$	1.0 $\times 10^7$
(Cp-Me ₅) ₂ Fe ^{+/0}	3.8 (\pm 0.4) $\times 10^7$ ^e	d	5.0 $\times 10^7$	2.0 $\times 10^7$
Cp ₂ Fe ^{+/0}	5.3 (\pm 0.8) $\times 10^6$ ^e	1.6 (\pm 0.6) $\times 10^6$ ^e	5.0 $\times 10^7$	2.0 $\times 10^7$
(C ₆ H ₆) ₂ Cr ^{+/0}	---	6.0 (\pm 0.7) $\times 10^7$ ^f	---	1.4 $\times 10^7$

^a Cp = cyclopentadienyl

^b Measured rate constant for homogeneous self exchange of listed couple in acetonitrile (ACN) or dimethylsulfoxide (DMSO), as indicated in column heading.

^c This work, obtained in 0.1M tetraethylammonium tetrafluoroborate (see footnote 7 for details of nmr spectra).

^d Value could not be obtained due to insufficient solubility of ferrocene

^e Values from ref. 3, obtained in absence of added electrolyte (values decrease by ca. 20% upon adding 0.1M electrolyte³).

^f Value from ref. 4, obtained in absence of added electrolyte.

^g Rate constant for homogeneous self exchange, derived from Eq. (1). Values of various parameters obtained as follows: K_p (0.26M^{-1}) from Eq. (2) with $r = 7.6\text{\AA}$, $\delta r = 0.6\text{\AA}$; $\kappa_{\text{el}} = 1$; ν_n from "overdamped" solvent relaxation model [Eq. (9) of ref. 1b]; ΔG_{os}^* from dielectric continuum model [Eq. (19) of ref. 1b]; ΔG_{is}^* from bond distance and force constant data, (see text), taken as follows: (Cp-Me₅)₂Co^{+/0}, 0.5 kcal mol⁻¹; Cp₂Co^{+/0}, 0.7 kcal mol⁻¹; (Cp-Me₅)₂Fe^{+/0} and Cp₂Fe^{+/0}, 0.35 kcal mol⁻¹; (C₆H₆)₂Cr^{+/0}, 0.5 kcal mol⁻¹.

END

FILMED

MARCH, 19 88

DTIC